

PHOTOCATALYTIC DECOLORIZATION OF TARTRAZINE USING PERIWINKLE SHELL ASH: EFFECT OF OPERATIONAL VARIABLES, KINETIC AND ISOTHERM STUDY

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ABSTRACT

The potential of a locally sourced material, periwinkle shell ash (PSA), has been explored as an effective photocatalyst for the photocatalytic decolorization of tartrazine in aqueous solution. The effect of operational variables such as irradiation time, initial dye concentration and catalyst loading on the extent of decolorization was investigated. The results show that the optimum levels of the variables were 50 minutes, 30 mg/L and 5 g/L for irradiation time, initial dye concentration and catalyst loading respectively. The pseudo first order and the Langmuir–Hinshelwood kinetic models were able to sufficiently describe the kinetics of the process. The diffusion mechanism was described by the intra-particle diffusion model, while the adsorption equilibrium was described by the Langmuir isotherm equation. The results obtained indicate that PSA can be used as an effective photocatalyst for the removal of tartrazine from aqueous solution.

Keywords: Periwinkle shell ash; tartrazine; kinetics; isotherm; equilibrium; azo dye

INTRODUCTION

Water pollution due to the discharge of colored effluents from textile dye manufacturing and textile dye mills is one of the major environmental concerns in the world today [1]. It has been reported that between 1 and 20% of the dye is lost during the dyeing process and it is subsequently released in wastewater [2]. The release of these colored wastewaters poses serious aesthetic and ecological problems such as eutrophication to the aquatic ecosystems [3]. Azo dyes are considered to be the largest group among the established classes of synthetic dyes and an important example is tartrazine. Tartrazine is a useful colorant widely employed in food products, drugs, cosmetics etc. Nevertheless, the toxic and harmful effects of tartrazine have been established by numerous studies [4-7]. Hence, these types of dyes have been classified as potentially carcinogenic and mutagenic by the US National Institution for Occupational Safety and Health [8]. Their chemical structures are characterized by highly substituted aromatic rings joined by one or more azo groups [9]. These substituted ring structures make the dye molecules recalcitrant and not readily amenable to degradation by conventional wastewater treatment methods [10].

Different physical, chemical and biological methods have been applied for the treatment of dye-containing effluents. Physical methods are not effective, they do not

degrade the dye and the dyes are merely transferred to another phase, thereby introducing the need for further treatment steps [11]. Biological methods appear to be effective and economical. However, this treatment does not lead to the mineralization of the dyes as they are only adsorbed on the sludge [3]. Heterogeneous photocatalytic degradation facilitated by semiconductor catalysts has been identified as a promising technique for the removal of most dyes from liquid effluents [10, 12]. The photocatalytic reaction takes place when the semiconductor particle absorbs a photon of light which is more energetic than its band gap, leading to the excitation of the electron from the valence band to the conduction band. A hole–electron pair is formed in the process and this initiates the oxidation of the dye [13]. The important advantage of this method is that it often results in complete mineralization and degradation of most pollutants that are not amenable to other treatment processes [14]. Furthermore, it is cheaper than most processes and it can be carried out under ambient conditions [1]. Several studies have been carried out on photocatalytic degradation of dyes, mainly with TiO₂/UV or ZnO/UV systems and these studies established the factors that influence the rate of the photocatalytic reaction [15-20]. However, TiO₂ has an important shortcoming in that it absorbs only a limited amount of photons in the UV region. Hence it requires a high-power UV excitation source to function effectively [21]. Some researchers have reported the preparation of photocatalysts from waste materials [22-25]. However, very little work has been reported on the preparation of photocatalysts from periwinkle shells [12, 26, 27]. Hence the focus of this work is to evaluate the potential application of periwinkle shell ash as photocatalyst for the photocatalytic decolorization of tartrazine in aqueous solutions. The effects of operational variables such as irradiation time, initial dye concentration and catalyst loading on the decolorization process were investigated. The photocatalytic decolorization of tartrazine was further evaluated by carrying out kinetic (pseudo first order, pseudo second order, intra particle diffusion and Langmuir–Hinshelwood models) and isotherm studies (Langmuir and Freundlich isotherms).

MATERIALS AND METHODS

Preparation and Characterization of Periwinkle Shell Ash

Periwinkle shells were sourced locally from Benin City in Edo State of Nigeria. The shells were washed and dried in an oven at 110 °C to constant mass, followed by crushing and calcination at 600 °C in a muffle furnace and subsequent sieving to obtain fine particles (< 350 µm) of periwinkle shell ash (PSA). The prepared PSA was characterized by determining the composition using X-ray fluorescence (XRF) analysis. X-ray diffraction (XRD) was used to determine the ultimate elemental composition of the PSA using a Philips X-ray diffractometer [12, 28]. The functional groups on the PSA surface were determined by Fourier transform infrared spectrometry (FTIR) using a Perkin Elmer spectrum 100 FT–IR spectrometer in the frequency range 4000 to 400 cm⁻¹, operating in ATR (attenuated total reflectance) mode. The surface structure and other properties of the PSA were evaluated by the nitrogen adsorption method at -196 °C. The surface area of the PSA was determined using the standard BET equation.

Dye

All reagents and chemicals used in this study were of analytical grade and were used without further purification. Tartrazine dye with molecular formula $C_{16}H_9N_4Na_3O_9S_2$ and molecular weight 534.36 g/mol was obtained from Stanvac Laboratories in Benin City, Edo State, Nigeria. The molecular structure of the dye is shown in Figure 1.

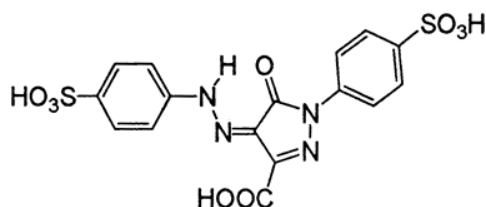


Figure 1. Chemical structure of tartrazine “trisodium 1-(4-sulfonatophenyl)-4-(4-sulfonatophenylazo)-5-pyrazolone-3-carboxylate” [8].

Photoreactor

The experimental setup has been described previously [12]. The reactor was irradiated by four UV lamps emitting at 254 nm (40 W, Philips TUV G6T5). The intensity of the light emitted by the lamps was measured as 1.65 mW/cm^2 using a Lux-UV-IR meter (Leybold Co.).

Photocatalytic Degradation Studies

The dye solution was prepared in appropriate concentrations using deionized water. The photocatalytic decolorization studies were conducted using 4000 mL of dye solution and a predetermined quantity of photocatalyst. With the UV lamps still off, air was bubbled at a predetermined flowrate through the dye solution in the holding tank. The lamps were subsequently turned on and the pump was started to convey the dye–catalyst suspension into the reactor. Samples of 25 mL aliquots were withdrawn at regular intervals and were immediately centrifuged at 8000 rpm for 10 min to remove suspended catalyst particles. The progress of the photocatalytic decolorization process was monitored by measuring the absorbance of the samples with a UV–Visible spectrophotometer (Shimadzu UV 2101 PC) at a wavelength of 520 nm. A calibration curve based on Beer–Lambert’s law was established by relating the absorbance to the concentration. The effect of irradiation time, initial dye concentration and photocatalyst loading on the decolorization efficiency was then investigated. The percentage photocatalytic decolorization of tartrazine dye was calculated using Eq. (1).

$$\text{Decolorization efficiency} = \frac{C_o - C_t}{C_o} \times 100 \quad (1)$$

The amount of tartrazine adsorbed at time t , (q_t) and at equilibrium (q_e) was calculated using Eqs. (2)–(3).

$$q_t = \frac{V_s(C_o - C_t)}{W} \quad (2)$$

$$q_e = \frac{V_s(C_o - C_e)}{W} \quad (3)$$

where C_o , C_e and C_t are the initial, equilibrium and instantaneous concentrations of tartrazine respectively. V_s is the volume of the aqueous solution and W is the amount of catalyst.

RESULTS AND DISCUSSION

Characterization of Photocatalyst

The results of the chemical composition analysis of the PSA have been previously reported [12, 26]. The XRF results indicated that the major constituents of the PSA were calcium oxide (CaO), silica (SiO₂) and aluminum oxide (Al₂O₃), which respectively accounted for 41.3, 33.2 and 9.2% of the weight of PSA analyzed. XRD results indicated that the major elements in PSA were iron (Fe) and zinc (Zn), which respectively accounted for 19.2 and 16.5% of the weight of PSA analyzed. Other oxides and elements were also present but in small proportions. Navaladian et al. [29] reported that transition metals and their oxides are known to exhibit catalytic properties, thus supporting the choice of PSA for this study. Similar observations have been reported by other researchers [18, 30]. The surface area, bulk density, and porosity of the PSA were similar to those reported in the literature [31, 32]. FTIR results revealed peaks of 3626.59 cm⁻¹, 1797.58 cm⁻¹, 1561.43 cm⁻¹ and 1374.34 cm⁻¹ in the infrared spectrum of periwinkle shell ash, corresponding to O–H, C=O, C=C and C–H bonds respectively. Absorption bands in the range of 2700 to 1430 cm⁻¹ and 900 to 1380 cm⁻¹ indicate the presence of phenyl groups and the stretching of the C–O bond in carboxylic groups present in the PSA [33]. These FTIR bands represent functional groups that possess strong bonds which can be protonated at slightly acidic conditions to be potential adsorption sites for organic molecules [34].

Effect of Irradiation Time

Figure 2 shows the effect of irradiation time on the photocatalytic decolorization of tartrazine by PSA. The decolorization efficiency showed a steady and progressive increase with increase in irradiation time to attain a maximum value of about 81% at an irradiation time of 50 minutes. Increasing the irradiation time beyond 50 minutes did not have any impact on the decolorization process. The initial increase in the decolorization efficiency could be attributed to the abundance of active sites on the surface of the photocatalyst, which favored the photocatalytic reaction [35]. As the reaction proceeds further, these active sites are then occupied by several layers of dye molecules such that the rate of the decolorization reaction slows down. Furthermore, the insignificant change in the decolorization efficiency observed beyond 50 minutes of irradiation time could also indicate that the system had attained a state of equilibrium, which means that the catalyst had become saturated and could no longer adsorb any more dye molecules [8]. Equilibrium times in the range of 90 to 120 minutes were reported by Chakrabarti and Dutta [36] under varying conditions of photocatalytic degradation of model textile dyes in wastewater using ZnO as photocatalyst. Akyol et al. [37] reported lower equilibrium times of 25 and 35 minutes respectively for the photocatalytic

decolorization of Remazol Red RR in aqueous systems using ZnO and TiO₂ as photocatalysts. These results show that the photocatalytic decolorization of dyes is a fast kinetic process.

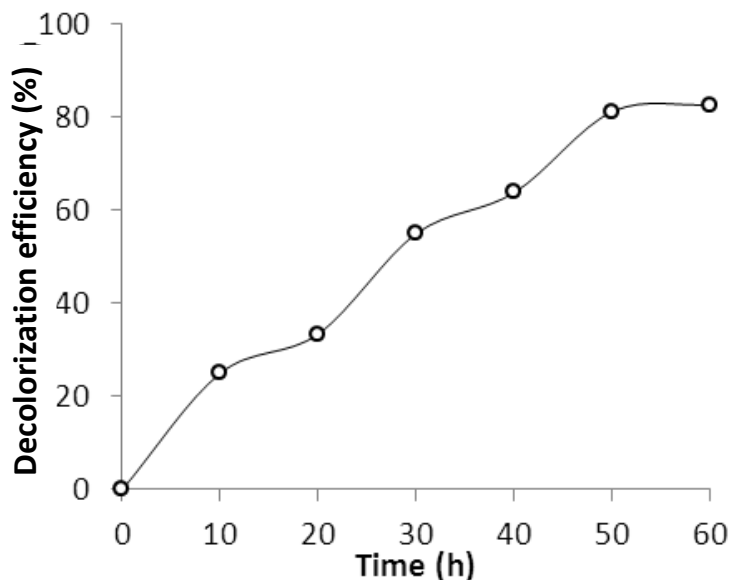


Figure 2. Effect of irradiation time on the photocatalytic decolorization of tartrazine (PSA loading, 5 g/L; initial concentration, 30 mg/L; temperature, 25 °C).

Effect of Initial Dye Concentration

The initial concentration of the dye solution influenced the decolorization process as shown in Figure 3. The decolorization efficiency initially increased from about 8% to about 83% when the initial dye concentration was increased from 10 to 30 mg/L. Increasing the concentration of dye beyond 30 mg/L resulted in a decrease in the decolorization efficiency. Konstantinou and Albanis [38] noted that the rate of dye degradation generally increases with the increase in dye concentration up to a certain level beyond which any further increase in dye concentration does not positively affect the degradation process. Similar observations have been reported by Sakthivel et al. [18]. The rate of the degradation reaction is dependent on the rate of formation of the hydroxyl radicals as well as the probability of the radicals reacting with the dye molecules. As the concentration of dye is increased, there will be more of the dye molecules in solution and this will increase the probability of the radicals reacting with the dye molecules, consequently resulting in an enhancement of the rate of the degradation of the dye [38]. However, as the concentration of dye is increased beyond the optimum point, the decolorization efficiency decreases because of the decrease in the formation of hydroxyl radicals on the surface of the catalyst. The reason for this could be that the active sites on the catalyst surface are occupied by layers of dye molecules [39, 40]. Furthermore, at high concentration of dye, a significant amount of the UV radiation could be absorbed by the dye molecules rather than the PSA particles and this reduces the efficiency of the decolorization process because of the reduction in the rate of formation of the oxidizing radicals [41].

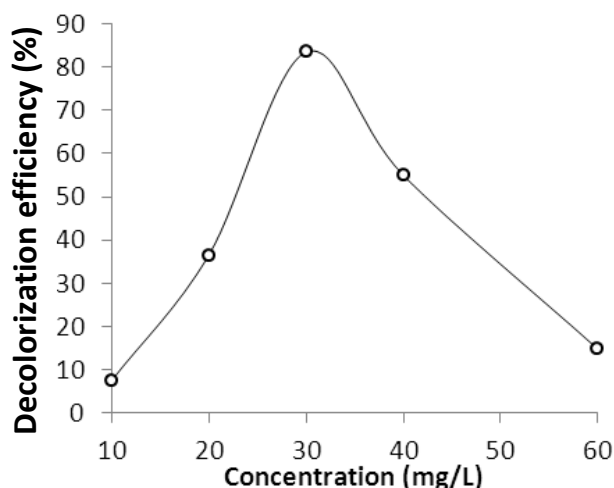


Figure 3. Effect of initial dye concentration on the photocatalytic decolorization of tartrazine (PSA loading, 5 g/L; temperature, 25 °C).

Effect of Catalyst Loading

Figure 4 shows the effect of PSA loading on the decolorization efficiency. With increase in PSA loading from 1 to 5 g/L, the decolorization efficiency increased from about 19 to 85%. Further increase in PSA loading did not have any positive impact on the decolorization process. The initial increase in decolorization efficiency observed might be rationalized in terms of the availability of active sites on the PSA surface. The availability of active sites increases with increase in catalyst loading, hence the rate of decolorization is enhanced [15].

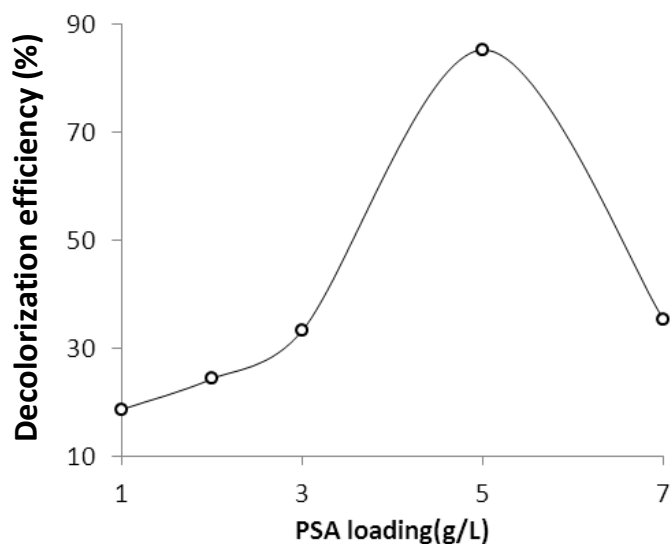


Figure 4. Effect of PSA loading on the photocatalytic decolorization of tartrazine (initial concentration, 30 mg/L; temperature, 25 °C).

Inasmuch as elevated levels of catalyst loading increase the availability of active sites on the catalyst surface, they also lead to a reduction in the light penetration and consequently a reduction in the photoactivated volume of the suspension. This explains

the decrease in decolorization efficiency observed beyond a catalyst loading of 5 g/L. Some researchers suggested that at high catalyst loading, agglomeration and sedimentation of the catalyst particles could occur, making a fraction of the catalyst surface unavailable for photon and dye absorption [42]. According to Suri et al. [43], the optimum catalyst dosage could be dependent on the initial solute concentration. They further suggested that it might not be advisable to use higher catalyst loadings as a result of possible aggregation and reduced irradiation field.

Kinetics of Photodegradation

The kinetics of the photodegradation process was studied using four kinetic models, namely pseudo first order, pseudo second order, intra particle diffusion and Langmuir–Hinshelwood kinetic models.

Pseudo First Order Model

The pseudo first order kinetic model is expressed in its integrated linear form as follows [44]:

$$\ln(q_e - q_t) = \ln q_e - k_1 t \tag{4}$$

where q_e and q_t (mg/g) are adsorption capacity at equilibrium and at time t respectively. k_1 is the rate constant of pseudo first order adsorption (1/min). The values of $\ln(q_e - q_t)$ were plotted against t . The plot of $\ln(q_e - q_t)$ against t resulted in a linear relationship from which the values of k_1 and q_e were determined as shown in Figure 5. The values of the constants are given in Table 1. The linear relationship observed in Figure 5 suggests that the pseudo first order model was applicable to the process.

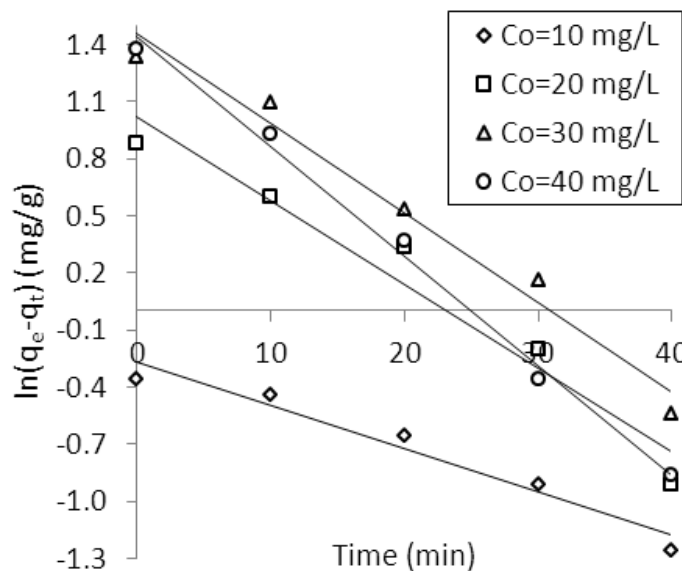


Figure 5. Pseudo first order model results for tartrazine photodegradation by PSA (PSA loading, 5 g/L; temperature, 25 °C).

Pseudo Second Order Model

The pseudo second order kinetic model is expressed in its integrated linear form as follows [45]:

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{1}{q_e} t \quad (5)$$

where k_2 is the rate constant of the pseudo second order process (g/mg.min). The plot of (t/q_t) versus t as shown in Figure 6 was used to calculate the values of the model parameters. The kinetic constants calculated from the plot at different initial dye concentrations are shown in Table 1. It is very obvious from Figure 6 that the pseudo second order model was not able to represent the kinetics of the degradation process, as seen in the very low R^2 values (Table 1).

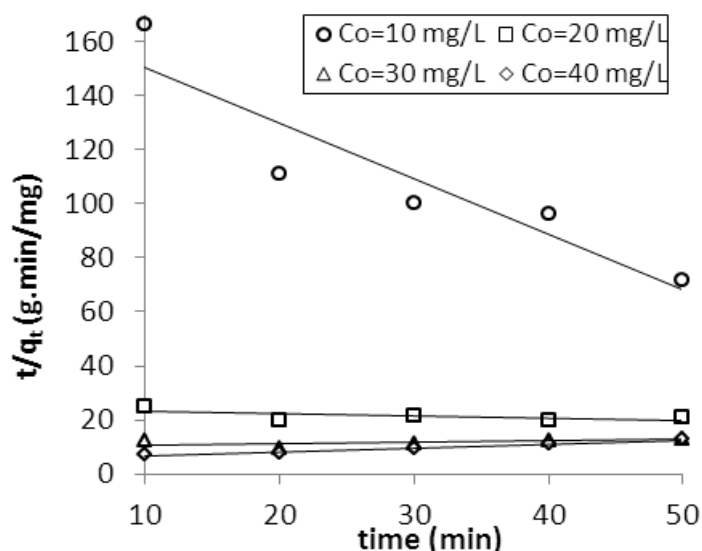


Figure 6. Pseudo second order model results for tartrazine photodegradation by PSA (PSA loading, 5 g/L; temperature, 25 °C).

Intra-Particle Diffusion Model

The diffusion mechanism of the process was modeled using the intra-particle diffusion model as follows [46]:

$$q_t = K_p t^{1/2} + C \quad (6)$$

where K_p is the intra-particle diffusion rate constant (mg/g.min^{1/2}) and C is a measure of the boundary layer effect. The values of the intra-particle diffusion model parameters at different initial dye concentrations are presented in Table 1. The plot of q_t versus $t^{1/2}$ from which the parameters were calculated is shown in Figure 7. The plot indicates the existence (although not significant) of some boundary layer effect and further showed that intra-particle diffusion was not the only rate-limiting step. The high R^2 values obtained for the kinetic parameters estimated within the concentration range

investigated showed that the intra-particle diffusion model was able to describe the diffusion mechanism of the process.

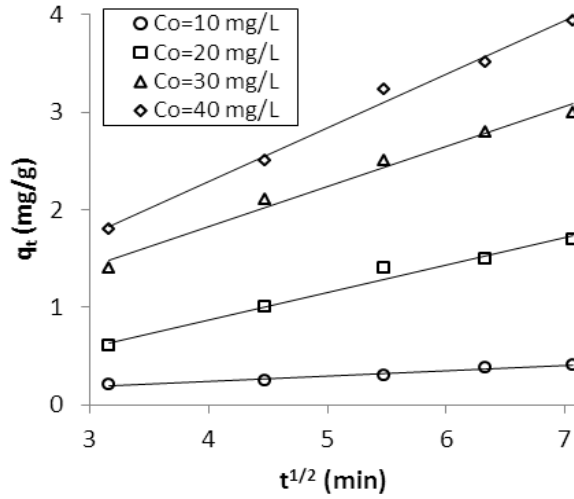


Figure 7. Intra-particle diffusion model results for tartrazine photodegradation by PSA (PSA loading, 5 g/L; temperature, 25 °C).

Langmuir–Hinshelwood Model

The Langmuir–Hinshelwood kinetic model developed by Turchi and Ollis [47] can be expressed as follows:

$$r_o = -\frac{dc}{dt} = \frac{k_r K C_{eq}}{1 + K C_{eq}} \tag{7}$$

where r_o is the initial rate of reaction in mg/L.min, k_r is the rate constant for photocatalysis in mg/L min, K is the rate constant for adsorption in L/mg, C_{eq} is the concentration of bulk solution in mg/L at adsorption equilibrium, c is the concentration of bulk solution at any time t . Eq. (7) may be expressed in its linearized form as follows:

$$\frac{1}{r_o} = \frac{1}{k_r K} \frac{1}{C_{eq}} + \frac{1}{k_r} \tag{8}$$

Values of $1/r_o$ were plotted against $1/C_{eq}$ for the dye as shown in Figure 8 and the values of the calculated constants are presented in Table 1.

Table 1. Kinetic constant parameter values for the photocatalytic decolorization of tartrazine.

C_o (mg/L)	Pseudo first order			Pseudo second order model			Intra-particle diffusion model			Langmuir–Hinshelwood model	
	k_f	q_e	R^2	k_2	q_e	R^2	K_p	C	R^2	k_r	
10	0.023	0.76	0.958	0.024	-0.49	0.846	0.05	0.0182	0.972	k_r	3.158
20	0.044	2.76	0.950	0.00029	-12.05	0.404	0.28	0.258	0.979	K	0.0066
30	0.047	4.28	0.976	0.00018	22.73	0.244	0.41	0.1905	0.984	R^2	0.960
40	0.058	4.22	0.994	0.004	6.85	0.982	0.55	0.0791	0.991		

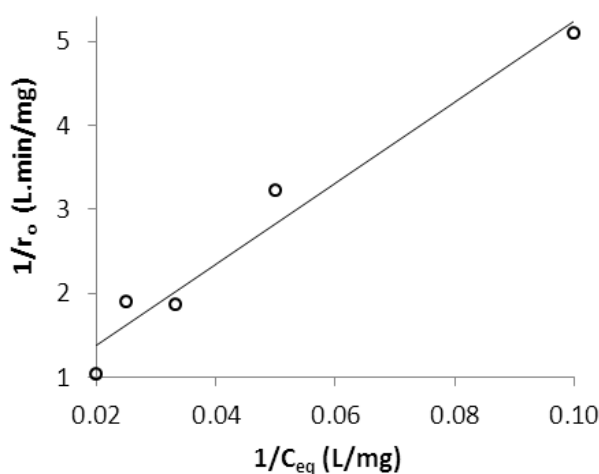


Figure 8. Langmuir–Hinshelwood model results for tartrazine photodegradation by PSA (PSA loading, 5 g/L; temperature, 25 °C).

Adsorption Isotherm Studies

The Langmuir and Freundlich adsorption isotherm models were used to analyze the equilibrium data for the photodegradation of tartrazine dye.

Langmuir Isotherm

The Langmuir isotherm model which assumes monolayer adsorption of the adsorbate can be expressed in its linear form as follows [48]:

$$\frac{C_e}{q_e} = \frac{1}{q_o} C_e + \frac{1}{K_L q_o} \quad (9)$$

where q_o is the maximum sorption capacity (mg/g) of the adsorbent, while K_L is the sorption constant (L/mg). A linear plot of C_e/q_e against C_e as shown in Figure 9 was used to calculate the values of q_o and K_L , which are presented in Table 2 alongside the correlation coefficient (R^2) value.

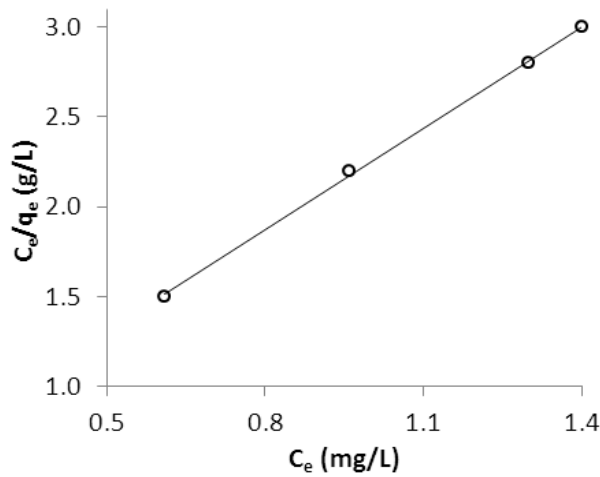


Figure 9. Langmuir isotherm model results for tartrazine photodegradation by PSA (PSA loading, 5 g/L; temperature, 25 °C).

Table 2. Kinetic parameters for Langmuir and Freundlich isotherms.

Langmuir isotherm			Freundlich isotherm		
q_o (mg/g)	K_L (L/mg)	R^2	K_f (mg/g)	N	R^2
0.531	5.160	0.999	0.393	1.560	0.959

The separation factor (R_L) can be used to determine whether or not the adsorption process will be favorable [49]. The parameter predicts that the type of isotherm could be irreversible ($R_L=0$), favorable ($0 < R_L < 1$) or unfavorable ($R_L > 1$).

$$R_L = \frac{1}{(1 + K_L C_o)} \tag{10}$$

C_o is the initial concentration of tartrazine.

Table 3. R_L values and type of isotherm.

Initial concentration (mg/L)	R_L Value
10	0.019
20	0.010
30	0.006
40	0.009

For this study, the values of R_L given in Table 3 are between zero and one, indicating that the adsorption was favorable.

Freundlich Isotherm

The Freundlich equation is expressed as follows:

$$q_e = K_f (C_e)^{1/n} \quad (11)$$

This equation can be expressed in its linearized form as follows:

$$\ln q_e = \ln K_f + 1/n \ln C_e \quad (12)$$

where K_f and n are the Freundlich constants related to the adsorption capacity and adsorption intensity respectively. A linear plot of $\ln q_e$ against $\ln C_e$ as shown in Figure 10 was employed to calculate the values of the constants. The values of these parameters as well as the correlation coefficient (R^2) of the Freundlich equation are given in Table 2. Values of n between 1 and 10 typically indicate beneficial and favorable adsorption. Chakrabarti and Dutta [36] reported K_f and n values of 0.431 and 1.343 for the photocatalytic degradation of Methylene Blue in wastewater using ZnO as semiconductor catalyst. Agarry and Aremu [50] suggested that the difference between the values could be as a result of differences in the range of concentration, type of material used, pH, temperature and properties of the adsorbent. The high values of the correlation coefficients obtained for the Langmuir isotherm as shown in Table 2 indicate that the data conformed well to the Langmuir isotherm equation.

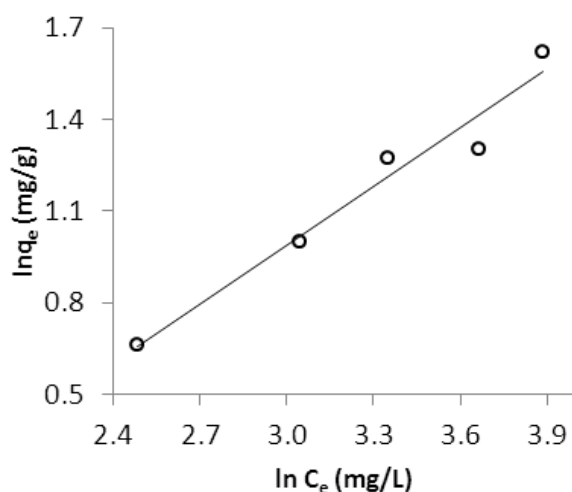


Figure 10. Freundlich isotherm model results for tartrazine photodegradation by PSA (PSA loading, 5 g/L; temperature, 25 °C).

CONCLUSIONS

Photocatalytic decolorization of tartrazine was carried out in a continuous flow reactor using periwinkle shell ash as photocatalyst. The decolorization process was affected by factors such as irradiation time, initial dye concentration and PSA loading dosage. The optimum values of irradiation time, initial dye concentration and PSA loading were 50 minutes, 30 mg/L and 5 g/L. The kinetics of photodegradation of tartrazine was well described by the pseudo first order and the Langmuir–Hinshelwood kinetic models with

high correlation coefficient values ($R^2 > 0.9$). The diffusion mechanism and adsorption equilibrium were well described by the intra-particle diffusion model and the Langmuir isotherm equation respectively. PSA has proven to be a suitable alternative to commercially available photocatalysts like TiO_2 and ZnO .

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REFERENCES

- [1] Pare B, Jonnalagadda S, Tomar H, Singh P, Bhagwat V. ZnO assisted photocatalytic degradation of acridine orange in aqueous solution using visible irradiation. *Desalination*. 2008;232:80-90.
- [2] Akpan U, Hameed B. Parameters affecting the photocatalytic degradation of dyes using TiO_2 -based photocatalysts: a review. *Journal of Hazardous Materials*. 2009;170:520-9.
- [3] Sauer T, Neto GC, Jose H, Moreira R. Kinetics of photocatalytic degradation of reactive dyes in a TiO_2 slurry reactor. *Journal of Photochemistry and Photobiology A: Chemistry*. 2002;149:147-54.
- [4] Gao Y, Li C, Yin H, An X, Jin H. Effect of food azo dye tartrazine on learning and memory functions in mice and rats, and the possible mechanisms involved. *Journal of Food Science*. 2011;76:T125-T9.
- [5] Hannuksela M HT. Food additive hypersensitivity-near myth. *Duodecim*. 2009;125:527-32.
- [6] Mehedi N, Mokrane N, Alami O, Ainad-Tabet S, Zaoui C, Kheroua O, et al. A thirteen week ad libitum administration toxicity study of tartrazine in Swiss mice. *African Journal of Biotechnology*. 2013;12:4519-29.
- [7] Mpountoukas P, Pantazaki A, Kostareli E, Christodoulou P, Kareli D, Poliliou S, et al. Cytogenetic evaluation and DNA interaction studies of the food colorants amaranth, erythrosine and tartrazine. *Food and Chemical Toxicology*. 2010;48:2934-44.
- [8] Behnajady M, Modirshahla N, Hamzavi R. Kinetic study on photocatalytic degradation of CI Acid Yellow 23 by ZnO photocatalyst. *Journal of Hazardous Materials*. 2006;133:226-32.
- [9] Forgacs E, Cserhati T, Oros G. Removal of synthetic dyes from wastewaters: a review. *Environment International*. 2004;30:953-71.
- [10] Bergamini RB, Azevedo EB, De Araújo LRR. Heterogeneous photocatalytic degradation of reactive dyes in aqueous TiO_2 suspensions: decolorization kinetics. *Chemical Engineering Journal*. 2009;149:215-20.
- [11] Chen X, Mao SS. Titanium dioxide nanomaterials: synthesis, properties, modifications, and applications. *Chemical Reviews*. 2007;107:2891-959.
- [12] Felix A, Amenaghawon A, Mededode A. Heterogeneous photocatalytic degradation of naphthalene using periwinkle shell ash: effect of operating variables, kinetic and isotherm study. *South African Journal of Chemical Engineering*. 2014;19:31-45.
- [13] Matilainen A, Sillanpää M. Removal of natural organic matter from drinking water by advanced oxidation processes. *Chemosphere*. 2010;80:351-65.

- [14] Akyol A, Bayramoglu M. The degradation of an azo dye in a batch slurry photocatalytic reactor. *Chemical Engineering and Processing: Process Intensification*. 2008;47:2150-6.
- [15] Inamdar J, Singh S. Photocatalytic detoxification method for zero effluent discharge in dairy industry: Effect of operational parameters. *International Journal of Chemical and Biological Engineering*. 2008;1:160-4.
- [16] RamMohan G, Rosenberger P, Buxy S, Pullammanappallil P, Goswami T. Titanium Dioxide Mediated Photocatalytic Degradation of Methylene Blue in a Fixed Film-Type Photoreactor. *AATCC Review*. 2014;14:40-4.
- [17] Sahel K, Bouhent M, Belkhadem F, Ferchichi M, Dappozze F, Guillard C, et al. Photocatalytic degradation of anionic and cationic dyes over TiO₂ P25, and Ti-pillared clays and Ag-doped Ti-pillared clays. *Applied Clay Science*. 2014;95:205-10.
- [18] Sakthivel S, Neppolian B, Shankar M, Arabindoo B, Palanichamy M, Murugesan V. Solar photocatalytic degradation of azo dye: comparison of photocatalytic efficiency of ZnO and TiO₂. *Solar Energy Materials and Solar Cells*. 2003;77:65-82.
- [19] Srinivasan A, Viraraghavan T. Decolorization of dye wastewaters by biosorbents: a review. *Journal of Environmental Management*. 2010;91:1915-29.
- [20] Zanon MVB, Sene JJ, Anderson MA. Photoelectrocatalytic degradation of Remazol Brilliant Orange 3R on titanium dioxide thin-film electrodes. *Journal of Photochemistry and Photobiology A: Chemistry*. 2003;157:55-63.
- [21] Liu G, Wang Z, Zheng W, Yang S, Sun C. Visible-Light-Driven Photocatalytic Degradation of Aniline over NaBiO₃. *Advances in Condensed Matter Physics*. 2014;2014.
- [22] Altın I, Sökmen M. Preparation of TiO₂-polystyrene photocatalyst from waste material and its usability for removal of various pollutants. *Applied Catalysis B: Environmental*. 2014;144:694-701.
- [23] da Silva WL, Lansarin MA, Stedile FC, dos Santos JH. The potential of chemical industrial and academic wastes as a source of supported photocatalysts. *Journal of Molecular Catalysis A: Chemical*. 2014;393:125-33.
- [24] Sapiña M, Jimenez-Relinque E, Castellote M. Turning waste into valuable resource: potential of electric arc furnace dust as photocatalytic material. *Environmental Science and Pollution Research*. 2014;21:12091-8.
- [25] Sugrañez R, Cruz-Yusta M, Mármol I, Morales J, Sánchez L. Preparation of sustainable photocatalytic materials through the valorization of industrial wastes. *ChemSusChem*. 2013;6:2340-7.
- [26] Aisien FA, Amenaghawon NA, Ekpenisi EF. Photocatalytic decolourisation of industrial wastewater from a soft drink company. *Journal of Engineering and Applied Sciences*. 2013;9:11-6.
- [27] Osarumwense JO AF. Application of local pozzolans in the photodegradation of toluene. *Nigerian Journal of Biomedical Engineering*. 2012;10:13-9.
- [28] Aku S, Yawas D, Madakson P, Amaren S. Characterization of periwinkle shell as asbestos-free brake pad materials. *The Pacific Journal of Science and Technology*. 2012;13:57-63.
- [29] Navaladian S, Janet C, Viswanathan B, Viswanath R. On the possible treatment procedures for organic contaminants. *Photo/electrochemistry & photobiology in the environment. Energy and Fuel*. 2007:1-51.

- [30] Palmisano G, Augugliaro V, Pagliaro M, Palmisano L. Photocatalysis: a promising route for 21st century organic chemistry. *Chemical Communications*. 2007;3425-37.
- [31] Owabor C, Iyaomolere A. Evaluation of the Influence of Salt Treatment on the Structure of Pyrolyzed Periwinkle Shell. *Journal of Applied Sciences and Environmental Management*. 2013;17:321-7.
- [32] Umoh AA, Olusola KO. Compressive Strength and Static Modulus of Elasticity of Periwinkle Shell Ash Blended Cement Concrete. *International Journal of Sustainable Construction Engineering and Technology*. 2012;3:45-55.
- [33] Park D, Yun Y-S, Park JM. Studies on hexavalent chromium biosorption by chemically-treated biomass of *Ecklonia* sp. *Chemosphere*. 2005;60:1356-64.
- [34] Royer B, Cardoso NF, Lima EC, Vaghetti JC, Simon NM, Calvete T, et al. Applications of Brazilian pine-fruit shell in natural and carbonized forms as adsorbents to removal of methylene blue from aqueous solutions—Kinetic and equilibrium study. *Journal of Hazardous Materials*. 2009;164:1213-22.
- [35] Umoren S, Etim U, Israel A. Adsorption of methylene blue from industrial effluent using poly (vinyl alcohol). *Journal of Materials and Environmental Science*. 2013;4:75-86.
- [36] Chakrabarti S, Dutta BK. Photocatalytic degradation of model textile dyes in wastewater using ZnO as semiconductor catalyst. *Journal of Hazardous Materials*. 2004;112:269-78.
- [37] Akyol A, Yatmaz H, Bayramoglu M. Photocatalytic decolorization of Remazol Red RR in aqueous ZnO suspensions. *Applied Catalysis B: Environmental*. 2004;54:19-24.
- [38] Konstantinou IK, Albanis TA. TiO₂-assisted photocatalytic degradation of azo dyes in aqueous solution: kinetic and mechanistic investigations: a review. *Applied Catalysis B: Environmental*. 2004;49:1-14.
- [39] Daneshvar N, Salari D, Khataee A. Photocatalytic degradation of azo dye acid red 14 in water: investigation of the effect of operational parameters. *Journal of Photochemistry and Photobiology A: Chemistry*. 2003;157:111-6.
- [40] Faramarzpour M, Vossoughi M, Borghei M. Photocatalytic degradation of furfural by titania nanoparticles in a floating-bed photoreactor. *Chemical Engineering Journal*. 2009;146:79-85.
- [41] Shanthi M, Kuzhalosai V. Photocatalytic degradation of an azo dye, Acid Red 27, in aqueous solution using nano ZnO. *Indian Journal of Chemistry-Part A InorganicPhysical Theoretical and Analytical*. 2012;51:428.
- [42] So C, Cheng MY, Yu J, Wong P. Degradation of azo dye Procion Red MX-5B by photocatalytic oxidation. *Chemosphere*. 2002;46:905-12.
- [43] Suri RP, Thornton HM, Muruganandham M. Disinfection of water using Pt-and Ag-doped TiO₂ photocatalysts. *Environmental Technology*. 2012;33:1651-9.
- [44] El Nemr A, El Sikaily A, Khaled A, Abdelwahab O. Removal of toxic chromium (VI) from aqueous solution by activated carbon using *Casuarina equisetifolia*. *Chemistry and Ecology*. 2007;23:119-29.
- [45] Ho YS. Review of second-order models for adsorption systems. *Journal of Hazardous Materials*. 2006;136:681-9.
- [46] Weber WJ, Morris JC. Kinetics of adsorption on carbon from solution. *Journal of the Sanitary Engineering Division*. 1963;89:31-60.

- [47] Turchi CS, Ollis DF. Photocatalytic degradation of organic water contaminants: mechanisms involving hydroxyl radical attack. *Journal of Catalysis*. 1990;122:178-92.
- [48] Langmuir I. The adsorption of gases on plane surfaces of glass, mica and platinum. *Journal of the American Chemical Society*. 1918;40:1361-403.
- [49] Annadurai G, Ling LY, Lee J. Biodegradation of phenol by *Pseudomonas pictorum* on immobilized with chitin. *African Journal of Biotechnology*. 2007;6.
- [50] Agarry S, Aremu M. Batch equilibrium and kinetic studies of simultaneous adsorption and biodegradation of phenol by pineapple peels immobilized *Pseudomonas aeruginosa* NCIB 950. *British Biotechnology Journal*. 2012;2:26.